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CHAPTER 12 SOLUTIONS MANUAL States of Matter Section 12.1 Gases pages 402–410 Practice Problems page 405 1. Calculate the ratio of effusion rates for nitrogen (N₂) and neon (Ne). $\frac{\text{Rate N}}{\text{Rate Ne}} = 0.849$ $\frac{\text{Rate N}}{\text{Rate Ne}} \sqrt{\frac{\text{molar mass Ne}}{\text{molar mass N}}} = \frac{0.849 \sqrt{20.18}}{\sqrt{28.02}}$ 2. Calculate the ratio of diffusion rates for carbon monoxide and carbon dioxide. $\frac{\text{Rate CO}}{\text{Rate CO}_2} = 1.25$

States of Matter

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4.2 For the reaction, $2A + B \rightarrow A_2B$, the rate = $k[A][B]^2$ with $k = 2.0 \times 10^{-6} \text{ mol}^{-2} \text{ L}^2 \text{ s}^{-1}$. Calculate the initial rate of the reaction when $[A] = 0.1 \text{ mol L}^{-1}$, $[B] = 0.2 \text{ mol L}^{-1}$. Calculate the rate of reaction after $[A]$ is reduced to 0.06 mol L^{-1} . Sol. Initial rate of reaction = $k[A][B]^2 = (2.0 \times 10^{-6} \text{ mol}^{-2} \text{ s}^{-1})(0.1 \text{ mol L}^{-1})(0.2 \text{ mol L}^{-1})^2 = 8 \times 10^{-9} \text{ mol L}^{-1} \text{ s}^{-1}$.

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