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Chapter 9 Stoichiometry Test Answer Key Modern Chemistry This quiz covers information from Chapter 9, Section 9-2 on Stoichiometry. You will be expected to solve mole-mole, mole-mass, mass-mole and mass-mass problems. This quiz can only be taken once. Upon

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space provided, write the letter of the
term or phrase that best completes each
statement or best answers each
question. 1. Knowing the mole ratio of a
reactant and product in a chemical
reaction would

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Chapter 9 - Stoichiometry 9-1

Introduction to Stoichiometry

Composition Stoichiometry - deals with mass relationships of elements in

compounds Reaction Stoichiometry -

Involves mass relationships between reactants and products in a chemical

reaction I. Reaction Stoichiometry

Problems A. Four problem Types, One Common Solution

Chapter 9 - Stoichiometry

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1. 2. Mole Ratios a. Mole ratios are conversion factors that relate the number of moles of one chemical to another.
0.0209 mol AgNO₃ Aspirin production a. 13.5 kg C₉H₈O₄ b. 7.66 kg C₄H₆O₃ c. 4.29 L HC₂H₃O₂ Ideal stoichiometry calculations do not account for...

Chapter 9 Review Stoichiometry Answer Key Modern Chemistry

CHAPTER 9 REVIEW Stoichiometry
MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided.
1. Given the following equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$
a. What is the value of the coefficient x in this equation? 4.07 g/mol
b. What is the molar mass of C₃H₄? 40.07 g/mol
c. What is the mole ratio of O₂ to H₂O? 2:1 mol O₂:1 mol H₂O

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mole ratio of O₂ to H

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Stoichiometry 4 Chapter 9 Assignment & Problem Set 5. Isopropyl alcohol (C₃H₇OH) burns in air according to this equation: $2\text{C}_3\text{H}_7\text{OH}(\text{l}) + 9\text{O}_2(\text{g}) \rightarrow 6\text{CO}_2(\text{g}) + 8\text{H}_2\text{O}(\text{g})$ a. Calculate the moles of oxygen needed to react with 3.40 mol C₃H₇OH. b. Find the moles of each product formed when 3.40 mol C₃H₇OH reacts with oxygen. General Stoichiometric Calculations

Chapter 9 Homework - Maine- Endwell Middle School

Chapter 9: Standard Review Worksheet

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1. Answers will vary. An example is included below: $2\text{H}_2\text{O}_2(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l}) + \text{O}_2(\text{g})$ This describes the decomposition reaction of hydrogen peroxide. Microscopic: Two molecules of hydrogen peroxide (in aqueous solution) decompose to produce two molecules of liquid water and one molecule of oxygen gas.

Chapter 9: Standard Review Worksheet

Chapter 9 focuses on reaction stoichiometry: using a balanced chemical equation to calculate the number of grams, moles, or particles of reactants/products involved in a chemical reaction. Students had an introduction to composition stoichiometry in Chapter 3 and will now move on to some more difficult problems.

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